Pt. 60, App. A-4, Meth. 7B

Analyte	CAS No.	Sensitivity
Nitrogen oxides (NO _x), as NO ₂ , including: Nitric oxide (NO)	10102–43–9	
Nitrogen dioxide (NO ₂)	10102–44–0	30-786 ppmv

- $1.2\,$ Applicability. This method is applicable for the determination of NO_X emissions from nitric acid plants.
- 1.3 Data Quality Objectives. Adherence to the requirements of this method will enhance the quality of the data obtained from air pollutant sampling methods.

2.0 Summary of Method

2.1 A grab sample is collected in an evacuated flask containing a dilute sulfuric acid-hydrogen peroxide absorbing solution; the $NO_{\rm X}$, excluding nitrous oxide (N₂O), are measured by ultraviolet spectrophotometry.

3.0 Definition. [Reserved]

4.0 Interferences. [Reserved]

5.0 Safety

- 5.1 This method may involve hazardous materials, operations, and equipment. This test method may not address all of the safety problems associated with its use. It is the responsibility of the user of this test method to establish appropriate safety and health practices and to determine the applicability of regulatory limitations prior to performing this test method.
- 5.2 Corrosive reagents. The following reagents are hazardous. Personal protective equipment and safe procedures are useful in preventing chemical splashes. If contact occurs, immediately flush with copious amounts of water at least 15 minutes. Remove clothing under shower and decontaminate. Treat residual chemical burn as thermal burn.
- 5.2.1 Hydrogen Peroxide (H_2O_2) . Irritating to eyes, skin, nose, and lungs.
- 5.2.2 Sodium Hydroxide (NaOH). Causes severe damage to eyes and skin. Inhalation causes irritation to nose, throat, and lungs. Reacts exothermically with limited amounts of water.
- 5.2.3 Sulfuric Acid (H_2SO_4). Rapidly destructive to body tissue. Will cause third degree burns. Eye damage may result in blindness. Inhalation may be fatal from spasm of the larynx, usually within 30 minutes. May cause lung tissue damage with edema. 3 mg/m³ will cause lung damage in uninitiated. 1 mg/m³ for 8 hours will cause lung damage or, in higher concentrations, death. Provide ventilation to limit inhalation. Reacts violently with metals and organics.

6.0 Equipment and Supplies

- 6.1 Sample Collection. Same as Method 7, Section 6.1.
- 6.2 Sample Recovery. The following items are required for sample recovery:
- 6.2.1 Wash Bottle. Polyethylene or glass.
- 6.2.2 Volumetric Flasks. 100-ml (one for each sample).
- 6.3 Analysis. The following items are required for analysis:
- 6.3.1 Volumetric Pipettes. 5-, 10-, 15-, and 20-ml to make standards and sample dilutions.
- 6.3.2 Volumetric Flasks. 1000- and 100-ml for preparing standards and dilution of samples.
- 6.3.3 Spectrophotometer. To measure ultraviolet absorbance at 210 nm.
- 6.3.4 Analytical Balance. To measure to within 0.1 mg.

7.0 Reagents and Standards

NOTE: Unless otherwise indicated, all reagents are to conform to the specifications established by the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available. Otherwise, use the best available grade.

- 7.1 Sample Collection. Same as Method 7, Section 7.1. It is important that the amount of hydrogen peroxide in the absorbing solution not be increased. Higher concentrations of peroxide may interfere with sample analysis.
- 7.2 Sample Recovery. Same as Method 7, Section 7.2.
- 7.3 Analysis. Same as Method 7, Sections 7.3.1, 7.3.3, and 7.3.4, with the addition of the following:
- 7.3.1 Working Standard KNO₃ Solution. Dilute 10 ml of the standard solution to 1000 ml with water. One milliliter of the working standard is equivalent to $10~\mu g$ NO₂.
- 8.0 Sample Collection, Preservation, Storage, and Transport
- 8.1 Sample Collection. Same as Method 7, Section 8.1.
- 8.2 Sample Recovery.
- 8.2.1 Let the flask sit for a minimum of 16 hours, and then shake the contents for 2 minutes.
- 8.2.2 Connect the flask to a mercury filled U-tube manometer. Open the valve from the flask to the manometer, and record the flask temperature (T_f) , the barometric pressure, and the difference between the mercury levels in the manometer. The absolute internal